



LESSON 10

READING & WRITING CHEMICAL FORMULAS

I. Calculating Percent Composition by Mass

A. Every _____ is made up of specific _____ in fixed, whole-number

1. The _____ of an element is _____
2. The _____ of each _____ found in a _____ is
always the _____

B. Percent composition by _____ of a substance can be used to find the
_____ of a _____

C. Steps to calculating percent composition

1. Find the _____ of the _____
2. Find the mass of each _____ in the _____
3. _____ the mass of the _____ by the mass of
the _____, and _____ by _____

II. Examples

A. What is the percent composition of each element in Na_2SO_4 ?

1. Mass of the entire compound



2. Na = _____

3. S = _____

4. O = _____

B. What is the percent composition of each element in $\text{Mg}(\text{OH})_2$?

1. Mass of the entire compound



2. Mg = _____

3. O = _____

4. H = _____

C. What is the percent composition of each element in AgNO_3 ?

1. Mass of the entire compound



2. Ag = _____

3. N = _____

4. O = _____

